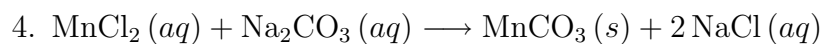
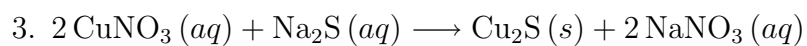
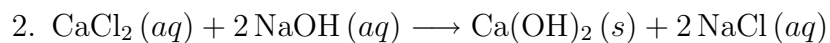
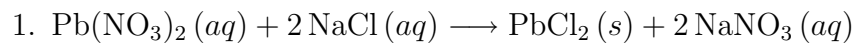


Name: _____

Honors Chemistry: yellow blue red

Net Ionic Equations and K_{sp}

For each of the following equations, write the net ionic equation and the expression for the solubility product constant K_{sp} .



Calculate the concentrations of the positive and negative ions in each of the following saturated solutions:

5. BaSO_4 , $K_{\text{sp}} = 1.1 \times 10^{-10}$

6. FeS , $K_{\text{sp}} = 6.0 \times 10^{-18}$

7. Al(OH)_3 , $K_{\text{sp}} = 4.6 \times 10^{-33}$

8. Hg_2Cl_2 , $K_{\text{sp}} = 1.3 \times 10^{-18}$